Naming and Equations Review

Name the following compounds. Remember, they may be either ionic or covalent compounds, so make sure you use the right naming method!

1) LiBr _____

2) Na_2O _____

3) $Ca(SO_4)$ _____

4) $Be(NO_3)_2$ _____

5) NO _____

Write the formulas for the following compounds. Remember, they may be either ionic or covalent compounds, so make sure you use the right method!

6) sodium nitride _____

7) sulfur dioxide _____

8) ammonia _____

9) calcium (phosphate) _____

10) aluminum (nitrate) _____

More Mixed Naming Fun!

Name these compounds. They may be either ionic or covalent.

1) Li(OH) _____

2) PBr_3 _____

3) $Na_2(SO_4)$ _____

4) $(NH_4)_2S$ _____

5) $Ca(CO_3)$ _____

6) CF_4 _____

7) $Na(NO_3)$ _____

8) P_2S_3 _____

9) $Al(NO_3)_3$ _____

10) $Mg(OH)_2$ _____

Write the formulas for the following compounds. Remember, they may be either ionic or covalent compounds, so make sure you use the right method!

11) potassium oxide _____

12) phosphorus tribromide _____

13) calcium (hydroxide) _____

14) dinitrogen sulfide _____

15) carbon monoxide _____

16) diboron tetrahydride _____

17) phosphorus pentabromide _____

18) sulfur dichloride _____

19) sodium (carbonate) _____

20) aluminum (acetate) _____

Naming and Equations Review – Answers

Name the following compounds. Remember, they may be either ionic or covalent compounds, so make sure you use the right naming method!

1) LiBr (Ionic) **lithium bromide**

2) Na_2O (Ionic) **sodium oxide**

3) $CaSO_4$ (Ionic) **calcium sulfate**

4) $Be(NO_3)_2$ (Ionic) **beryllium nitrate**

5) NO (covalent) **nitrogen monoxide**

Write the formulas for the following compounds. Remember, they may be either ionic or covalent compounds, so make sure you use the right method!

6) sodium nitride (ionic) $\mathbf{Na_3^{+1}N^{-3}}$

7) sulfur dioxide (covalent) $\mathbf{S^{+4}O_2^{-2}}$

8) ammonia (covalent) $\mathbf{N^{+3}H_3^{-1}}$

9) calcium phosphate (ionic) $\mathbf{Ca_3^{+2}(PO_4^{-3})_2}$

10) aluminum nitrate (ionic) $\mathbf{Al^{+3}(NO_3^{-1})_3}$

More Mixed Naming Fun! - Answers

Name these compounds. They may be either ionic or covalent.

1) LiOH (ionic) **lithium hydroxide**

2) PBr_3 (covalent) **phosphorus tribromide**

3) Na_2SO_4 (ionic) **sodium sulfate**

4) $(NH_4)_2S$ (ionic) **ammonium sulfide**

5) $CaCO_3$ (ionic) **calcium carbonate**

6) CF_4 (covalent) **carbon tetrafluoride**

7) $NaNO_3$ (ionic) **sodium nitrate**

8) P_2S_3 (covalent) **diphosphorus trisulfide**

9) $Al(NO_3)_3$ (ionic) **aluminum nitrate**

10) $Mg(OH)_2$ (ionic) **magnesium hydroxide**

Write the formulas for the following compounds. Remember, they may be either ionic or covalent compounds, so make sure you use the right method!

11)	potassium oxide	(ionic) $\mathbf{K_2^{+1}O^{-2}}$
12)	phosphorus tribromide	(covalent) $\mathbf{P^{+3}Br^{-1}_3}$
13)	calcium hydroxide	(ionic) $\mathbf{Ca^{+2}(OH^{-1})_2}$
14)	dinitrogen sulfide	(covalent) $\mathbf{N^{+1}_2S^{-2}}$
15)	carbon monoxide	(covalent) $\mathbf{C^{+2}O^{-2}}$
16)	diboron tetrahydride	(covalent) $\mathbf{B^{+2}_2H^{-1}_4}$
17)	phosphorus pentabromide	(covalent) $\mathbf{P^{+5}Br^{-1}_5}$
18)	sulfur dichloride	(covalent) $\mathbf{S^{+2}Cl^{-1}_2}$
19)	sodium carbonate	(ionic) $\mathbf{Na^{+1}_2CO_3^{-2}}$
20)	aluminum acetate	(ionic) $\mathbf{Al^{+3}(C_2H_3O_2^{-1})_3}$