**Guided Reading Chapter 16 Section 2**

1. All compounds have an electrical charge of \_\_\_\_\_\_\_\_\_.
2. one b) two c) zero
3. An oxidation number is the quantity that indicates the charge on an atom when it has gained, lost or \_\_\_\_\_\_\_\_\_\_\_\_ electrons.
4. Copy figure 16.12 on page 361, showing oxidation numbers of some common elements.
5. Would Beryllium tend to lose two electrons or gain six when forming bonds?
6. What is the most common oxidation number for group three on the Periodic table?
7. Elements near the noble gases tend to form \_\_\_\_\_\_\_\_\_\_\_ bonds.
8. ionic b) covalent c) metallic
9. The farther apart elements are on the Periodic Table the more likely they are to form \_\_\_\_\_\_\_\_\_\_ bonds.
10. ionic b) covalent c) metallic
11. Nonmetals tend to form \_\_\_\_\_\_\_\_\_\_\_ bonds.
12. Ionic b) covalent c) metallic
13. Using figure 16.14 on page 364, how many Chlorine atoms are needed to bond with a Copper (II) atom to form a compound?
14. What is a binary compound?
15. How many atoms of each element is in CaCO3?
16. What type of ion is one that contains more than one atom?
17. What is the oxidation number for peroxide?
18. How do you write the name of a binary compound?
19. How do you name a compound with polyatomic ions?